An estimation of the $^{18}\text{O}/^{16}\text{O}$ ratio of UT/LMS ozone based on artefact CO in air sampled during CARIBIC flights

S. Gromov¹, C. A. M. Brenninkmeijer¹
¹ Max Planck Institute for Chemistry, Mainz, Germany
Correspondence to: S. Gromov (sergey.gromov@mpic.de)

Abstract

An issue of O₃-driven artefact production of CO in the upper troposphere/lowermost stratosphere (UT/LMS) air analysed in the CARIBIC−1 project is being discussed. By confronting the CO mixing and isotope ratios obtained from different analytical instrumentation, we (i) reject natural/artificial sampling and mixing effects as possible culprits of the problem, (ii) ascertain the chemical nature and quantify the strength of the contamination, and (iii) demonstrate successful application of the isotope mass-balance calculations for inferring the isotope composition of the contamination source. The $\delta^{18}\text{O}$ values of the latter indicate the oxygen very likely being inherited from O₃. The $\delta^{13}\text{C}$ values hint at reactions of trace amounts of organics with stratospheric O₃ that could have yielded the artificial CO. While the exact contamination mechanism is not known, it is clear that the issue pertains only to the earlier (first) phase of the CARIBIC project. Finally, estimated UT/LMS ozone $\delta^{18}\text{O}$ values are lower than those observed in the stratosphere within the same temperature range, suggesting that higher pressures (240−270 hPa) imply lower isotope fractionation controlling the local $\delta^{18}\text{O}(\text{O}_3)$ value.

1 Introduction

[1] Accurate determination of the atmospheric carbon monoxide (CO) content based on the collection of air samples depends on the preservation of the mixing ratio of CO inside the receptacle, from the point of sampling to the moment of physicochemical analysis in a laboratory. A well known example in our field of research is the filling of pairs of glass flasks at South Pole
Station for analysis at NOAA in Boulder, Colorado, USA (Novelli et al., 1998). There, the duplicate air sampling allowed for a degree of quality control which in view of the long transit times, especially during polar winter, was a perhaps not perfect, but certainly a practical measure. Here we deal with a different case: Using aircraft-based collection of very large air samples rendered duplicate sampling unpractical, yet analyses could be performed soon after the sampling had taken place because of the proximity of the aircraft’s landing location to the laboratory involved. A presumption of the analytical integrity of the process was that the growth of CO in receptacles is gradual and takes its time. Reminding Thomas Henry Huxley’s statement, “The great tragedy of Science – the slaying of a beautiful hypothesis by an ugly fact”, it turned out, however, that for air we collected in stainless steel tanks in the upper troposphere/lowermost stratosphere (UT/LMS) higher CO values were measured in the laboratory than measured in situ during the collection of these air samples. Moreover, measurement of the stable oxygen isotopic composition of CO from these tanks revealed additional isotopic enrichments in $^{18}$O of 10% or more. It was soon realised that this phenomenon was due to the formation of CO in these tanks and/or possibly in the sampling system and inlet tubing used, by reactions involving ozone (Brenninkmeijer et al., 1999).

[2] Unexpectedly high $^{18}$O/$^{16}$O ratios in stratospheric ozone (O$_3$) were discovered by Konrad Mauersberger using a balloon-borne mass spectrometer (Mauersberger, 1981), which has triggered a series of theoretical and experimental studies on atmospheric O$_3$ heavy isotope enrichments (see, e.g., Schinke et al. (2006) for a review). In view of the advances in theoretical and laboratory studies on the isotopic composition of O$_3$ atmospheric measurements are welcome, they do however form a challenge. In the stratosphere O$_3$ number concentrations are high, but the remoteness of the sampling domain is a problem. In the troposphere, low O$_3$ number densities are the main obstacle, as indicated by few experiments performed to date (Krankowsky et al., 1995; Johnston and Thiemens, 1997; Vicars and Savarino, 2014). Nevertheless, recent analytical improvements, namely the use of an indirect method of reacting atmospheric O$_3$ with a substrate that can be analysed for the isotopic composition of the O$_3$-derived oxygen (Vicars et al., 2012), has greatly improved our ability to obtain information on the O$_3$ isotopic composition.

[3] Although the increase of CO concentrations in air stored in vessels is a well recognised problem, to our knowledge a specific O$_3$-related process has not been reported yet. Here we discuss this phenomenon and turn its disadvantage into an advantage, namely that of obtaining an estimate of the oxygen isotopic composition of O$_3$ in the UT/LMS, an atmospheric domain not
yet covered by specific measurements. The air samples we examine in this study were collected onboard a passenger aircraft carrying an airfreight container with analytical and air/aerosol sampling equipment on long distance flights from Germany to South India and the Caribbean within the framework of the CARIBIC project (Civil Aircraft for the Regular Investigation of the atmosphere Based on an Instrument Container, http://www.caribic-atmospheric.com).

2 Experimental and results

2.1 Whole air sampling

[4] CARIBIC−1 (Phase #1, abbreviated hereafter “C1”) was operational from November 1998 until April 2002 using a Boeing 767-300 ER operated by LTU International Airlines (Brenninkmeijer et al., 1999). Using a whole air sample (WAS) collection system, twelve air samples were collected per flight (of 8−10 hours duration at cruise altitudes of 10−12 km) in stainless steel tanks for subsequent laboratory analysis of the mixing ratios (i.e. mole fractions) of various trace gases, including $^{14}$CO. Large air samples were required in view of the ultra-low number density of this mainly cosmogenic tracer (10−100 molecules cm$^{-3}$ standard temperature and pressure (STP), about 0.4−4 amol/mol). Hereinafter STP denotes dry air at 273.15 K, 101325 Pa. Each C1 WAS sample (holding 350 litres of air STP) was collected over 15−20 min intervals representing the number density-weighted average of the compositions encountered along flight segments of about 250 km. The overall uncertainty of the measured WAS CO is less than ±1% for the mixing ratio and ±0.1‰/±0.2‰ for δ$^{13}$C(CO)/δ$^{18}$O(CO), respectively (Brenninkmeijer, 1993; Brenninkmeijer et al., 2001). Isotope compositions are reported throughout this manuscript using the so-called delta value δ = \( \frac{R}{R_{st}} - 1 \) relating the ratio \( R \) of rare ($^{13}$C, $^{18}$O or $^{17}$O) over abundant isotopes of interest to the standard ratio \( R_{st} \). These are Vienna Standard Mean Ocean Water (VSMOW) for $^{18}$O/$^{16}$O (Gonfiantini, 1978; Coplen, 1994) and $^{17}$O/$^{16}$O (Assonov and Brenninkmeijer, 2003), and Vienna Pee Dee Belemnite (VPDB) for $^{13}$C/$^{12}$C (Craig, 1957), respectively. As we mention above, the oxygen isotope composition of the CO present in these WAS samples was corrupted, in particular when O$_3$ levels were as high as 100−600 nmol/mol.

[5] CARIBIC−2 (Phase #2, referred to as “C2”) started operation in December 2004 with a Lufthansa Airbus A340-600 fitted with a new inlet system and air sampling lines, including per-fluoroalkoxy alkane (PFA) lined tubing for trace gas intake (Brenninkmeijer et al., 2007). No flask CO mixing/isotope ratio measurements are performed in C2.
2.2 On-line instrumentation

In addition to the WAS collection systems, both C1 and C2 measurement setups include different instrumentation for on-line detection of [CO] and [O3] (hereinafter the squared brackets denote the mixing ratio of the respective species). In situ CO analysis in C1 is done using a gas chromatography (GC)-reducing gas analyser which provides measurements every 130 s with an uncertainty of ±3 nmol/mol (Zahn et al., 2000). In C2, a vacuum ultraviolet fluorescence (VUV) instrument with lower measurement uncertainty and higher temporal resolution of ±2 nmol/mol in 2 s (Scharffe et al., 2012) is employed. Furthermore, the detection frequency for O3 mixing ratios has also increased, viz., from 0.06 Hz in C1 to 5 Hz in C2 (Zahn et al., 2002; Zahn et al., 2012).

2.3 Results

When comparing the CO mixing ratios in relation to those of O3 for C1 and C2, differences are apparent in the LMS, where C2 [CO] values are systematically lower. This is illustrated in Fig. 1 (a) which presents the LMS CO-O3 distribution of the C2 in situ measurements overlaid with the C1 in situ and WAS data. The entire C1 CO/O3 dataset is presented in Fig. 2. For the in situ CO datasets we calculated the statistics (Fig. 1 (b)) of the samples with respective O3 mixing ratios clustered in 20 nmol/mol bins, i.e. the median and spread of [CO] as a function of [O3] analysed. The interquartile range, IQR, is used in the current analysis as a robust measure of the data spread instead of the standard deviation. The LMS data exhibit large [CO] variations for [O3] between 300 and 400 nmol/mol, which primarily reflect pronounced seasonal variations in the NH tropospheric CO mixing ratio. With increasing [O3], [CO] decreases to typical stratospheric values, and its spread reduces to mere 3.5 nmol/mol and less, as [O3] surpasses 500 nmol/mol. Despite the comparable spread in C1 and C2 [CO], from 400 nmol/mol of [O3] onwards the C1 CO mixing ratios start to level off, with no samples below 35 nmol/mol having been detected, whereas the C2 levels continuously decline. By the 570–590 nmol/mol O3 bin, C1 [CO] of $39.7^{+0.7}_{-1.3}$ nmol/mol contains some extra 14 nmol/mol compared to $25.6^{+1.2}_{-1.1}$ nmol/mol typical for C2 values. Overall, at [O3] above 400 nmol/mol the conspicuously high [CO] is marked in about 200 in situ C1 samples, of which 158 and 69 emerge as statistically significant mild and extreme outliers, respectively, when compared against the number of C2 samples ($n > 3 \times 10^5$). The conventions here follow Natrelle (2003), i.e. ±1.5 and ±3 IQR ranges define the inner and outer statistical fences (ranges outside which the data points are considered mild and extreme outliers) of the C2 [CO] distribution in every O3 bin, respectively. The statistics include
the samples in bins with average [O$_3$] of 420–620 nmol/mol. None of C1 CO at [O$_3$] above 560 nmol/mol agrees with the C2 observations. Because the CO-O$_3$ distribution cannot have changed over the period in question, we find that an apparent relative excess CO of up to 55% justifies and investigation into sampling artefacts and calibration issues.

[8] Unnatural elevations in $\delta^{18}$O(CO) from WAS measurements are also evident, as shown in Figs. 3 and 4. The large $\delta^{18}$O(CO) elevations that reach beyond $+16\%$ are found to be proportional to the concomitant O$_3$ mixing ratios (denoted with colour) and are more prominent at lower [CO]. Lower $\delta^{18}$O(CO) values, however, are expected based on our knowledge of UT/LMS CO sources (plus their isotope signatures) and available in situ observations (Fig. 3, shown with triangles), as elucidated by Brenninkmeijer et al. (1996) (hereafter denoted as “B96”). That is, the greater the proportion of stratospheric CO, the greater its fraction stemming from methane oxidation with a characteristic $\delta^{18}$O of 0‰ or lower (Brenninkmeijer and Röckmann, 1997). This occurs because the CO sink at ruling UT/LMS temperatures proceeds more readily than its production, as the reaction of hydroxyl radical (OH) with CO, being primarily pressure-dependent, is faster than the temperature-sensitive reaction of OH with CH$_4$. Furthermore, as the lifetime of CO quickly decreases with altitude, transport-mixing effects take the lead in determining the vertical distributions of [CO] and $\delta^{18}$O(CO) above the tropopause, hence their mutual relationship. This is seen from the B96 data at [CO] below 50 nmol/mol that line-up in a near linear relationship towards the end-members with lowest $^{18}$O/$^{16}$O ratios. These result from the largest share of the $^{18}$O-depleted photochemical component and extra depletion caused by the preferential removal of C$^{18}$O in reaction with OH (fractionation about $+11\%$ at pressures below 300 hPa, Stevens et al., 1980; Röckmann et al., 1998b).

[9] We are confident that the enhancements of C1 C$^{18}$O originate from O$_3$, whose large enrichment in $^{18}$O (above $+60\%$ in $\delta^{18}$O, Brenninkmeijer et al., 2003) is typical and found transferred to other atmospheric compounds (see Savarino and Morin (2012) for a review). In Fig. 3 it is also notable that not only the LMS compositions are affected but elevations of (3−10)$\%$ from the bulk $\delta^{18}$O(CO) values are present in more tropospheric samples with [CO] of up to 100 nmol/mol. These result from the dilution of the least affected CO-rich tropospheric air by CO-poor, however substantially contaminated, stratospheric air, sampled into the same WAS tank. Such sampling-induced mixing renders an unambiguous determination of the artefact source’ isotope signature rather difficult, because neither mixing nor isotope ratios of the admixed air portions are known sufficiently well (see below).
Differences between the WAS and in situ measured [CO] – a possible indication that the δ^{18}O(CO) contamination pertains specifically to the WAS data – average at Δ(WAS–in situ) = (5.3±0.2) nmol/mol (±1 standard deviation of the mean, n = 408) and happen to be random with respect to any operational parameter or measured characteristic in C1, i.e. irrespective of CO or O_{3} abundances. The above mentioned discrepancy remained after several calibrations between the two systems had been performed, and likely results from the differences in the detection methods, drifts of the calibration standards used (see details in Brenninkmeijer et al., 2001) and a short-term production of CO in the stainless steel tanks during sampling. The large spread of Δ(WAS–in situ) of ±3.5 nmol/mol (±1σ of the population) ensues from the fact that the in situ sampled air corresponds to (2−4)% of the concomitantly sampled WAS volume, as typically 6−7 in situ collections of 5 s were made throughout one tank collection of 17−21 min. The integrity of the WAS CO is further affirmed by the unsystematic distribution of the artefact compositions among tanks (in contrast to that for δ^{18}O(CO_{2}) in C1 discussed by Assonov et al., 2009). Overall, the WAS and in situ measured CO mixing ratios correlate extremely well (adj. R^{2} = 0.972, slope of 0.992±0.008 (±1σ), n = 408). However, both anomalies in [CO] and δ^{18}O(CO) manifest clear but complex influences of the concomitant [O_{3}]. That is, the C1 in situ and WAS [CO] and δ^{18}O(CO) data very likely evidence artefacts pertaining to the same O_{3}-driven effect. Below we discuss and quantify these influences.

3 Discussion

Three factors may lead to the (artefact) distributions seen for C1 in situ [CO] at LMS O_{3} mixing ratios, namely:

(i) Strong (linear) natural mixing, such as enhanced stratosphere-troposphere exchange (STE), when a [CO] outside the statistically expected range results from the integration of air having dissimilar ratios of the tracers’ mixing ratios, viz. [O_{3}]:[CO]. For example, mixing of two air parcels in a 16%:84% proportion (by moles of air) with typical [O_{3}]:[CO] of 700:24 (stratospheric) and 60:125 (tropospheric), respectively, yields an integrated composition with [O_{3}]:[CO] of 598:40 which indeed corresponds to C1 data (this case is exemplified by the mixing curve in Fig. 1). Nonetheless, occurrences of rather high stratospheric CO mixing ratios (in our case, 40 nmol/mol at the concomitant [O_{3}] of 500−600 nmol/mol compared to the typical 24−26 nmol/mol) are rare. For instance, a deep STE similar to that described by Pan et al. (2004) was observed by C2 only once (cf. the outliers at [O_{3}] of 500 nmol/mol in Fig. 1), whereas the C1 outliers were exclusively registered in some 12 flights during
1997–2001. No relation between these outliers and the large-scale [CO] perturbation due to extensive biomass burning in 1997/1998 (Novelli et al., 2003) is established, otherwise elevated CO mixing ratios should manifest themselves at lower [O₃] as well. Other tracers detected in CARIBIC provide supporting evidence against such strongly STE-mixed air having been captured by C1. That is, the binned distributions for water vapour and de-trended N₂O mixing ratios (not shown here) are similar for C1 and C2. Whereas the small relative variations in atmospheric [N₂O] merely confirm matching [O₃] distributions in CARIBIC, the stratospheric [H₂O] distributions witness no [O₃]:[H₂O] values corresponding to those of the C1 outliers, suggesting the latter being unnaturally low.

(ii) Mixing effects can also occur artificially, originating from sampling peculiarities or data processing. Since the CARIBIC platform is not stationary, about 5 s long sampling of an *in situ* air probe in C1 implies integration of the air compositions encountered along some hundred metres, owing to the high aircraft speed. This distance may cover a transect between tropospheric and stratospheric filaments of different compositions. The effect of such ‘translational mixing’ can be simulated by averaging the sampling data with higher temporal frequency over longer time intervals. In this respect, the substantially more frequent CO data in C2 (sampling interval <1 s) were artificially averaged over a set of increasing intervals to reckon whether the long sampling period in C1 could be the culprit for skewing its CO–O₃ distribution. As a result, the original C2 data and their averages (equivalent to the C1 CO sample injection time) differ negligibly, as do the respective [O₃]:[CO] values. Our simulations of the ‘translational mixing’ effects confirm that the actual C2 CO–O₃ distribution in the region of interest ([O₃] of 540–620 nmol/mol) remains insensitive to averaging intervals of up to 300 s. Furthermore, a very strong artificial mixing with an averaging interval of at least 1200 s (comparable to C1 WAS sampling time) is required to yield the averages from the C2 data with [O₃]:[CO] characteristic for the C1 outliers.

(iii) In view of the above, it is unlikely that any natural or artificial mixing processes are involved in the stratospheric [CO] discrepancies seen in C1. We therefore conclude that the sample contamination in C1 occurred prior to the probed air reaching the analytical instrumentation and WAS sampling tanks in the container, since clearly elevated stratospheric CO mixing ratios are common to WAS and *in situ* data. Two more indications, viz. growing [CO] discrepancy with increasing O₃ abundance, and the strong concomitant signal in δ¹⁸O(CO), suggest that O₃-mediated production of CO took place. Further, by confronting the C1 and C2 [CO] measure-
ments in a regression analysis (detailed in Appendix A), we quantify the artefact component 
$
\text{CO}_c \text{ being chiefly a function of } \text{O}_3 \text{ mixing ratio as }
\]
\[
[\text{CO}_c] = b \cdot [\text{O}_3]^2, \quad b = (5.19 \pm 0.12) \cdot 10^{-5} \text{ [mol/nmol]},
\]
which is equivalent to 8–18 nmol/mol throughout the respective [O$_3$] range of 
400–620 nmol/mol (see Fig. 1 (d)). Subtracting this artefact signal yields the corrected \textit{in situ} 
C1 CO–O$_3$ distribution conforming to that of C2 (\textit{cf.} red symbols in Fig. 1 (a)). 

[15] Importantly, since we can quantify the contamination strength using only the O$_3$ mixing ra-
tio, the continuous \textit{in situ} C1 [O$_3$] data allow estimating the integral artefact CO component in 
each WAS sample and, if the isotope ratio of contaminating O$_3$ is known, to derive the initial 
$\delta^{18}$O(CO). The latter, as it was mentioned above, is subject to strong sample-mixing effects, 
which is witnessed by $\delta^{18}$O(CO) outliers even at relatively high [CO] up to 100 nmol/mol. Ac-
counting for such cases is, however, problematic since it is necessary to distinguish the propor-
tions of the least modified (tropospheric) and significantly affected (stratospheric) components 
in the resultant WAS sample mix. Since this information is not available, we applied an \textit{ad hoc} 
correction approach, as described in the following. This approach is capable of determining the 
contamination source (\textit{i.e.}, O$_3$) isotope signature as well.

3.1 Contamination isotope signatures

[16] We use the differential mixing model (MM, originally known as the “Keeling-plot”) in 
combination with the parameterisation of the artefact CO component (Eq. (1)) to derive the iso-
topic composition of the latter. This approach makes no assumptions on the isotope signatures 
of CO in the air portions mixed in a given WAS tank. The MM parameterises the admixing of 
the portion of artefact CO to the WAS sample with the "true" initial composition, as formulated 
below:

\[
[\text{CO}] = [\text{CO}_t] + [\text{CO}_c], \quad (2)
\]

\[
\delta(\text{CO})[\text{CO}] = \delta(\text{CO}_t)[\text{CO}_t] + \delta(\text{CO}_c)[\text{CO}_c], \quad (3)
\]

where indices $c$ and $t$ distinguish the components pertaining to the estimated contamination and 
“true” composition sought (\textit{i.e.}, [CO$_t$] and $\delta$(CO$_t$)), respectively. Here the contamination 
strength [CO$_c$] is derived by integrating Eq. (1) using the \textit{in situ} C1 [O$_3$] data for each WAS 
sample. By rewriting the above equation with respect to the isotope signature of the analysed 
CO, one obtains:

\[
\delta(\text{CO}) = \delta(\text{CO}_c) + (\delta(\text{CO}_t) - \delta(\text{CO}_c))[\text{CO}_t]/[\text{CO}], \quad (4)
\]
which signifies that linear regression of $\delta$(CO) as a function of the reciprocal of [CO] yields the estimated contamination signature $\delta$(CO) at ([CO])$^{-1}$ → 0 when invariable "true" compositions ([CO], $\delta$(CO)) are taken (the Keeling plot detailing these calculations is shown in Fig. 5). We therefore apply the MM described by Eq. (4) to the subsets of samples picked according to the same reckoned [CO] (within a ±2 nmol/mol window, $n > 7$). Such selection, however, may be insufficient: Due to the strong sampling effects in the WAS samples (see previous Section), it is possible to encounter samples that integrate different air masses to the same [CO] but rather different average $\delta$(CO). The solution in this case is to refer to the goodness of the MM regression fit, because the $R^2$ intrinsically measures the linearity of the regressed data, i.e. closeness of the “true” values in a regarded subset of samples, irrespective of underlying reasons for that.

[17] Higher $R^2$ values thus imply higher consistency of the estimate, as demonstrated in Fig. 6 showing the calculated $\delta$(CO) for [CO] below 80 nmol/mol as a function of the regression $R^2$. The latter decreases with greater [CO] (i.e., larger sample subset size, since tropospheric air is more often encountered) and, correspondingly, larger variations in $\delta$(CO). Ultimately, at lower $R^2$ the inferred $\delta^{18}$O(CO) converge to values slightly above zero expected for uncorrelated data, i.e. C1 $\delta^{18}$O(CO) tropospheric average. A similar relationship is seen for the $\delta^{13}$C(CO) values (they converge around −28‰), however, there are no consistent estimates found ($R^2$ is generally below 0.4). Since such is not the case for $\delta^{18}$O, the MM is not sufficiently sensitive to the changes caused by the contamination, which implies that the artefact CO $\delta^{13}$C should be within the range of the “true” $\delta^{13}$C(CO) values. Interestingly, the MM is rather responsive to the growing fraction of the CH$_4$-derived component in CO with increasing [O$_3$], as the $\delta^{13}$C(CO) value of −(47.2±5.8)‰ inferred at $R^2$ above 0.4 is characteristic for the $\delta^{13}$C of methane in the UT/LMS. It is important to note that we have accounted for the biases in the analysed C1 WAS $\delta^{13}$C(CO) expected from the mass-independent isotope composition of O$_3$ (see details in Appendix B).

[18] We derive the “best-guess” estimate of the admixed CO $^{18}$O signature at $\delta^{18}$O(CO) = +(92.0±8.3)‰, which agrees with the other MM results obtained at $R^2$ above 0.75. Taking the same subsets of samples, the concomitant $^{13}$C signature matches $\delta^{13}$C(CO) = −(23.3±8.6)‰, indeed at the upper end of the expected LMS $\delta^{13}$C(CO) variations of −(25–31)‰. Because of that, the MM is likely insensitive to the changes in $\delta^{13}$C(CO) caused by the contamination (the corresponding $R^2$ values are below 0.1). Upon the correction using the inferred $\delta^{18}$O(CO) value, the C1 WAS $\delta^{18}$O(CO) data agree with B96 (shown with red symbols in Fig. 3). That is, variations in the observed C$^{18}$O are driven by (i) the seasonal/regional changes in the composi-
tion of tropospheric air and by (ii) the degree of mixing or replacement of the latter with the
stratospheric component that is less variable in $^{18}$O. This is seen as stretching of the scattered
tropospheric values ([CO] above 60 nmol/mol) towards $\delta^{18}$O(CO) of around $-10\%$ at [CO] of
25 nmol/mol, respectively. The corrected C1 $\delta^{13}$C(CO) data (shown in Fig. 7) are found to be in
a $\pm 1\%$ agreement with the observations by B96, except for several deep stratospheric samples
([CO] below 40 nmol/mol). The latter were encountered during “ozone hole” conditions and
carried extremely low $\delta^{13}$C(CO) values, which was attributed to the reaction of methane with
available free Cl radicals (Brennkinkmeijer et al., 1996).

3.2 Estimate of $\delta^{18}$O(O$_3$)

[19] The contamination $^{18}$O signature inferred here ($\delta^{18}$O(CO$_c$) = $(92.0\pm8.3)\%$) likely pertains
to O$_3$ and is comparable to $\delta^{18}$O(O$_3$) values measured in the stratosphere at temperatures about
30 K lower than those encountered in the UT/LMS by C1 (see Table 1 for comparison). If no
other factors are involved (see below), this discrepancy in $\delta^{18}$O(O$_3$) should be attributed to the
local conditions, i.e. the higher pressures (typically 240–270 hPa for C1 cruising altitudes) at
which O$_3$ was formed. Indeed, the molecular lifetime (the period through which the species’
isotope reservoir becomes entirely renewed, as opposed to the “bulk” lifetime) of O$_3$ encoun-
tered along the C1 flight routes is estimated on the order of minutes to hours at daylight
(H. Riede, Max Planck Institute for Chemistry, 2010), thus the isotope composition of the pho-
tochemically regenerated O$_3$ resets quickly according to the local conditions. Virtual absence of
sinks, in turn, leads to “freezing” of the $\delta^{18}$O(O$_3$) value during night in the UT/LMS. Verifying
the current $\delta^{18}$O(O$_3$) estimate against the kinetic data, in contrast to the stratospheric cases, is
problematic. The laboratory studies on O$_3$ formation to date have scrutinised the concomitant
kinetic isotope effects (KIEs) as a function of temperature at only low pressures (67 mbar); the
attenuation of the KIEs with increasing pressure was studied only at room temperatures (see
Table 1, also Brenninkmeijer et al. (2003) for references). A rather crude attempt may be under-
taken by assuming that the formation KIEs become attenuated at higher pressures in a similar
(proportional) fashion to that measured at 320 K, however applied to the nominal low-pressure
values reckoned at (220–230) K. A decrease in $\delta^{15}$O(O$_3$) of about $(6–8)\%$ is expected from
such calculation (cf. last row in Table 1), yet accounting for a mere one-half of the $(13–15)\%$
discrepancy between the stratospheric $\delta^{18}$O(O$_3$) values and $\delta^{18}$O(CO$_c$).

[20] Lower $\delta^{18}$O(CO$_c$) values could result from possible isotope fractionation accompanying the
production of the artefact CO. Although not quantifiable here, oxygen KIEs in the O$_3 \rightarrow$ CO
conversion chain cannot be ruled out, recalling that the intermediate reaction steps are not identi-
fiable and the artefact CO represents at most 4% of all O₃ molecules. Furthermore, the yield
λₒ of CO from O₃ may be lower than unity (see details in Appendix A). On the other hand, the
inference that the contamination strength primarily depends on [O₃] indicates that the kinetic
fractionation may have greater effect on the carbon isotope ratios of the artefact CO produced
(the δ¹³C(CO₂) values) in contrast to the oxygen ones. That is because all reactive oxygen avail-
able from O₃ becomes converted to CO, whilst the concomitant carbon atoms are drawn from a
virtually unlimited pool whose apparent isotope composition is altered by the magnitude of the
¹³C KIEs.

[21] Besides KIEs, selectivity in the transfer of O atoms from O₃ to CO affects the resulting
δ¹⁸O(CO₂) value. The terminal O atoms in O₃ are enriched with respect to the molecular (bulk)
O₃ composition when the latter is above +70‰ in δ¹⁸O (Janssen, 2005; Bhattacharya et al., 2008), therefore an incorporation of only central O atoms into the artefact CO mole-
cules should result in a reduced apparent δ¹⁸O(CO₂) value. Such exclusive selection is, howev-
er, less likely from the kinetic standpoint and was not observed in available laboratory studies
(see Savarino et al. (2008) for a review). For instance, Röckmann et al. (1998a) established the
evidence of direct O transfer from O₃ to the CO produced in alkene ozonolysis. A reanalysis of
their results (in light of findings of Bhattacharya et al. (2008)) suggests that usually the terminal
atoms of the O₃ molecule become transferred (their ratio over the central ones changes from the
bulk 2:1 to 1:0 for various species). Considering the alternatives of the O transfer in our case
(listed additionally in Table 1), the equiprobable incorporation of the terminal and central O₃
atoms into CO should result in the δ¹⁸O(O₃) value in agreement with the “crude” estimate based
on laboratory data given above.

[22] Furthermore, the conditions that supported the reaction of O₃ (or its derivatives) followed by
the production of CO are vague. A few hypotheses ought to be scrutinised here. First, a fast
O₃ → CO conversion must have occurred, owing to short (i.e., fraction of a second) exposure
time of the probed air to the contamination. Accounting for the typical C1 air sampling condi-
tions (these are: sampled air pressure of 240–270 hPa and temperature of 220–235 K outboard
to 275–300 K inboard, sampling rate of 12.85·10⁻³ mol s⁻¹ corresponding to 350 L STP sam-
pled in 1200 s, inlet/tubing volume gauged to yield exposure times of 0.01 to 0.1 s due to vari-
able air intake rate, [O₃] of 600 nmol/mol), the overall reaction rate coefficient (kₑ in Eq. (A3)
from Appendix A) must be on the order of (6·10⁻¹⁵/τₑ) molecules⁻¹ cm³, where τₑ is the exposure
time. Assuming the case of a gas-phase CO production from a recombining O₃ derivative and
an unknown carbonaceous compound X, the reaction rate coefficient for the latter \((k\) in Eq. (A2)
in Appendix A) must be unrealistically high, at least \(6 \times 10^{-10}\) molec\(^{-1}\) cm\(^3\) s\(^{-1}\) over \(τ_c = 1/100\) s.
This number decreases proportionally with growing \(τ_c\) and \([X]\), if we take less strict exposure
conditions. Nonetheless, in order to provide the amounts of artefact CO we detect, a minimum
mixing ratio of 20 nmol/mol (or up to 4 µg of C per flight) of X is required, which is not availa-
ble in the UT/LMS from the species readily undergoing ozonolysis, \textit{e.g.} alkenes.

[23] Second, a more complex heterogeneous chemistry on the inner surface of the inlet or sup-
plying tubing may be involved. Such can be the tracers’ surface adsorption, (catalytic) decom-
position of O\(_3\) and its reaction with organics or with surface carbon that also may lead to the
production of CO (Oyama, 2000). Evidence exists for the dissociative adsorption of O\(_3\) on the
surfaces with subsequent production of the reactive atomic oxygen species (see, \textit{e.g.},
Li \textit{et al.}, 1998, also Oyama, 2000). It is probable that sufficient amounts of organics have re-
mained on the walls of the sampling line exposed to highly polluted tropospheric air, to be later
broken down by the products of the heterogeneous decomposition of the ample stratospheric O\(_3\).
Unfortunately, the scope for a detailed quantification of intricate surface effects in the C1 CO
contamination problem is very limited.

4 Conclusions

[24] Recapitulating, the \textit{in situ} measurements of CO and O\(_3\) allowed us to unambiguously quanti-
fy the artefact CO production from O\(_3\) likely in the sample line of the CARIBIC−1 instrumenta-
tion. Strong evidence to that is provided by the isotope CO measurements. We demonstrate the
ability of the simple mixing model (“Keeling-plot” approach) to single out the contamination
isotope signatures even in the case of a large sampling-induced mixing of the air with very dif-
ferent compositions. Obtained as a collateral result, the estimate of the δ\(^{18}\)O(O\(_3\)) in the UT/LMS
appears adequate, calling, however, for additional laboratory data (\textit{e.g.}, the temperature-driven
variations of the O\(_3\) formation KIE at pressures above 100 hPa) for a more unambiguous verifi-
cation.

Appendix A. Contamination assessment

[25] We quantify the C1 CO contamination strength (denoted [CO\(_c\)], obtained by discriminating
the C1 outliers from respective C2 data) in a sequence of regression analyses. We foremost as-
certain that no other species or operational parameter (\textit{e.g.} temperature, pressure, flight dura-
tion, season, latitude, time of day, etc.) measured in C1 appear to determine (e.g., systematically correlate with) \([\text{CO}_2]\), except that for \([\text{O}_3]\). We hypothesise therefore that a production of artefact CO molecules was initiated by \(\text{O}_3\) (via either its decomposition or a reaction with an unknown educt) and proceeded with incorporation of carbon (donated by some carbonaceous species \(X\)) and oxygen (donated by \(\text{O}_3\) or its derivatives) atoms into final CO. Despite that neither the actual reaction chain nor its intermediates are known, it is possible to describe the artefact component \(\text{CO}_c\) produced (hereinafter curly brackets \{} denote number densities) as

\[
\{\text{CO}_c\} = \lambda_{\text{O}_3} v \tau_c, \quad (A1)
\]

where the yield \(\lambda_{\text{O}_3}\), a diagnostic quantity, relates the amount of artefact CO molecules produced to the total number of \(\text{O}_3\) molecules consumed in the system, \(\tau_c\) denotes the reaction time (period throughout which sampled air is exposed to contamination), and \(v\) stands for the overall rate of the reaction chain. The latter, being regarded macroscopically (empirically), is parameterised to account for the order of reaction chain rate with respect to hypothesised reactants (McNaught and Wilkinson, 1997) as

\[
v = k \{X\}^\kappa \{\text{O}_3\}^\kappa, \quad (A2)
\]

where \(\kappa\) and \(K\) are the partial orders with respect to \(X\) and \(\text{O}_3\) number densities, respectively, and \(k\) is the rate coefficient. Here it is implied that changes to \(\{X\}\) and \(\{\text{O}_3\}\) are negligible throughout the exposure time \(\tau_c\) (typically < 0.1 s for C1 sample line). As stated above, we find that variations in \(\{\text{CO}_c\}\) correlate exclusively with variations in \(\{\text{O}_3\}\), hence Eq. (A2) can be reduced by assuming constancy of \(\{X\}\) and \(K\) to:

\[
v_c = k_c \{\text{O}_3\}^\kappa. \quad (A3)
\]

Here, \(k_c = k \{X\}^\kappa\) (often referred to as pseudo-first-order or “observed” rate coefficient) quantifies the rate of reaction chain exclusively propelled by \(\text{O}_3\). Finally, using Eqs. (A1) and (A3), the artefact \(\{\text{CO}_c\}\) component is expressed as

\[
\{\text{CO}_c\} = b \cdot \{\text{O}_3\}^\kappa, \quad b = \lambda_{\text{O}_3} k_c \tau_c \quad (A4)
\]

where the constant proportionality factor \(b\) integrates the influence of the unknown (and as we explicate below, likely invariable) \(\{X\}\), \(k\), \(K\) and \(\tau_c\).

[26] Eq. (A4) defines the regression expression using which we attempt to fit the values of \(\{\text{CO}_c\}\) as a function of \(\kappa\), \(\{\text{O}_3\}\) and \(b\). In the first regression iteration we keep both \(\kappa\) and \(b\) as free parameters, which provides best approximation at \(\kappa = 2.06 \pm 0.38\), suggesting reactions of two \(\text{O}_3\) molecules in case elementary reactions constitute the reaction mechanism, or two elementary steps involving \(\text{O}_3\) or its derivatives in case a stepwise reaction is involved.

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(McNaught and Wilkinson, 1997). In a subsequent regression iteration we set $\kappa = 2$, which yields better (as opposed to the first iteration) estimate of $b$ of $(5.19\pm0.12)\times10^{-5}$ mol/nmol ($\pm1\sigma$, adj. $R^2 = 0.83$, red. $\chi^2 = 4.0$; here the value of $b$ in mole fraction units is derived using the air density at C1 sampling conditions for relating fitted $[\text{CO}_c]$ and observed $[\text{O}_3]$). At last, we ascertain that the best regression results are obtained particularly at $\kappa = 2$, as indicated by the regression statistic ($R^2$ and $\chi^2$) that asymptotically improves when a set of regressions with neighbouring (i.e. below and above 2) integer values of $\kappa$ is compared. The low uncertainty (within $\pm3\%$) associated with the estimate of $b$ confirms an exclusive dependence of the contamination source on the O3 mixing ratio, as well as much similar reaction times $\tau_c$. The regressed value of $[\text{CO}_c]$ as a function of $[\text{O}_3]$ is presented in Fig. 1 (d) (solid line). It is possible to constrain the overall yield $\lambda_{\text{O}_3}$ of CO molecules in the artefact source chain to be between 0.5 and 1, comparing the magnitude of $[\text{CO}_c]$ to the discrepancy between the $[\text{O}_3]$ measured in C1 and C2 ($\pm20$ nmol/mol, taken equal to the $[\text{O}_3]$ bin size owing to the N2O−O3 and H2O−O3 distributions matching well between the datasets). Lower $\lambda_{\text{O}_3}$ values, otherwise, should have resulted in a noticeable (i.e., greater than 20 nmol/mol) decrease in the C1 O3 mixing ratios with respect to the C2 levels.

Appendix B. Corrections to measured $\delta^{13}\text{C}(\text{CO})$ values due to the oxygen MIF

[27] Atmospheric O3 carries an anomalous isotope composition (or mass-independent fractionation, MIF) with a substantially higher relative enrichment in $^{17}\text{O}$ over that in $^{18}\text{O}$ (above $+25\%$ in $\Delta^{17}\text{O} = (\delta^{17}\text{O}+1)/(\delta^{18}\text{O}+1)\beta−1$, $\beta = 0.528$) when compared to the majority of terrestrial oxygen reservoirs that are mass-dependently fractionated (i.e., with $\Delta^{17}\text{O}$ of $0\%$) (see Brenninkmeijer et al. (2003) and refs. therein). CO itself also has an unusual oxygen isotopic composition, possessing a moderate tropospheric MIF of around $+5\%$ in $\Delta^{17}\text{O}(\text{CO})$ induced by the sink KIEs in reaction of CO with OH (Röckmann et al., 1998b; Röckmann et al., 2002) and a minor source effect from the ozonolysis of alkenes (Röckmann et al., 1998a; Gromov et al., 2010). A substantial contamination of CO by O3 oxygen induces proportional changes to $\Delta^{17}\text{O}(\text{CO})$ that largely exceed its natural atmospheric variation. On the other hand, the MIF has implications in the analytical determination of $\delta^{13}\text{C}(\text{CO})$, because the presence of C17O species interferes with the mass-spectrometric measurement of the abundances of C13CO possessing the same basic molecular mass ($m/z$ is 45). When inferring the exact C17O/C18O ratio in the analysed sample is not possible, analytical techniques usually involve assumptions (e.g., mass-dependently fractionated
compositions or a certain non-zero $\Delta^{17}$O value) with respect to the C$^{17}$O abundances (Assonov and Brenninkmeijer, 2001). In effect for the C1 CO data, the artefact CO produced from O$_3$ had contributed with unexpectedly high C$^{17}$O abundances that led to the overestimated $\delta^{13}$C(CO) analysed. The respective bias $^{13}\delta_b$ is quantified using

$$^{13}\delta_b = 7.26 \cdot 10^{-2} \Delta^{17}$O(CO), \hspace{1cm} (B1)$$

where the actual $\Delta^{17}$O(CO) value is approximated from the natural CO MIF signal $^{17}\Delta_n$ and the typical O$_3$ MIF composition $^{17}\Delta_c$ as

$$\Delta^{17}$O(CO) = $(^{17}\Delta_n([CO] - [CO_c]) + ^{17}\Delta_c [CO_c])/[CO])^{-1}. \hspace{1cm} (B2)$$

Here [CO] and [CO$_c$] denote the analysed CO mixing ratio and contamination magnitude, respectively, used in the contamination assessment (see Appendix A, Eq. (A4)) and in calculations with the MM (see Sect. 3.1). For the purpose of the current estimate it is sufficient to take $^{17}\Delta_n$ of +5‰ representing equilibrium enrichments expected in the remote free troposphere and UT/LMS. For the O$_3$ MIF signature $^{17}\Delta_c$, the value of +30‰ (the average $\Delta^{17}$O(O$_3$) expected from the kinetic laboratory data at conditions met along the C1 flight routes, see Sect. 3.2 and Table 1) is adopted. The coefficient that proportionates $^{13}\delta_b$ and $\Delta^{17}$O in Eq. (B1) is derived by linearly regressing the $\delta^{13}$C(CO) biases (simulated using the calculation apparatus detailed by Assonov and Brenninkmeijer, 2001) as a function of $\Delta^{17}$O(CO) varying within a (0−30)‰ range for the CO with initially unaccounted MIF (e.g., the sample is assumed to be mass-dependently fractionated). It therefore quantifies some extra +(0.726±0.003)‰ in the analysed $\delta^{13}$C(CO) per every +10‰ of $\Delta^{17}$O(CO) excess. The most contaminated C1 WAS CO samples at [O$_3$] above 300 nmol/mol are estimated to bear $\Delta^{17}$O(CO) of (6−12)‰ corresponding to fractions of (0.10−0.27) of the artefact CO in the sample. Accordingly, the reckoned $\delta^{13}$C(CO) biases span (0.5−0.9)‰. Although not large, these well exceed the $\delta^{13}$C(CO) measurement precision of ±0.1‰ and were corrected for, and therefore are taken into account in the calculations with the MM presented in Sect. 3.1.

**Acknowledgements**

[28] The authors are indebted to Claus Koeppel, Dieter Scharffe and Dr. Andreas Zahn for their work and expertise on the carbon monoxide and ozone measurements in C1 and C2. Hella Riede is acknowledged for comprehensive estimates of the species lifetimes along the CARIBIC flight routes. We are grateful to Dr. Taku Umezawa, Dr. Angela K. Baker, Dr. Em-
ma C. Leedham, Dr. Sergey Assonov, the anonymous reviewer and Dr. Jan Kaiser for the helpful discussions and comments on the manuscript.

References


Waibel, A., Wandel, A., Waschitschek, K., Wiedensohler, A., Xueref-Remy, I., Zahn, A.,
Zech, U., and Ziereis, H.: Civil Aircraft for the regular investigation of the atmosphere based on an
instrumented container: The new CARIBIC system, *Atmos. Chem. Phys.*, 7, 4953–4976,

Coplen, T. B.: Reporting of stable hydrogen, carbon, and oxygen isotopic abundances (Technical Report),

Craig, H.: Isotopic standards for carbon and oxygen and correction factors for mass-spectrometric analysis
1957.


and its application to modelling the stable isotopic composition of atmospheric trace gases,

Guenther, J., Erbacher, B., Krankowsky, D., and Mauersberger, K.: Pressure dependence of two relative

coefficients and isotopologue fractionation in $^{16}$O–$^{18}$O oxygen mixtures, *Chem. Phys. Lett.*, 367,

Janssen, C.: Intramolecular isotope distribution in heavy ozone ($^{16}$O$^{18}$O$^{16}$O and $^{16}$O$^{16}$O$^{18}$O),

Johnston, J. C. and Thiemens, M. H.: The isotopic composition of tropospheric ozone in three


Krankowsky, D., Lämmerzahl, P., Mauersberger, K., Janssen, C., Tuzson, B., and Röckmann, T.:
Stratospheric ozone isotope fractionations derived from collected samples, *J. Geophys. Res. Atm.*, 112,

Li, W., Gibbs, G. V., and Oyama, S. T.: Mechanism of Ozone Decomposition on a Manganese Oxide
Catalyst. 1. In Situ Raman Spectroscopy and Ab Initio Molecular Orbital


McNaught, A. D. and Wilkinson, A.: IUPAC. Compendium of Chemical Terminology (the "Gold Book"),
XML on-line corrected version: http://goldbook.iupac.org (2006-) created by

Natrella, M.: NIST/SEMATECH e-Handbook of Statistical Methodosed., edited by: Croarkin, C. and
Tobias, P., NIST/SEMATECH, http://www.itl.nist.gov/div898/handbook/ (last access: 07 May 2014),
2003.

Novelli, P. C., Masarie, K. A., and Lang, P. M.: Distributions and recent changes of carbon monoxide in


**Tables**

**Table 1. Ozone \(^{18}\text{O}/^{16}\text{O}\) isotope ratios from literature and this study**

<table>
<thead>
<tr>
<th>Domain</th>
<th>T (K)</th>
<th>P (hPa)</th>
<th>(\delta^{18}\text{O}(\text{O}_3)) (%)</th>
<th>Remarks</th>
</tr>
</thead>
<tbody>
<tr>
<td>Stratosphere</td>
<td>190–210</td>
<td>13–50</td>
<td>83–93 (&lt;3)</td>
<td>1</td>
</tr>
<tr>
<td>UT/LMS</td>
<td>220–235</td>
<td>240–270</td>
<td>89–95 (8)</td>
<td>2</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td>84–88 (6) T</td>
<td></td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td>91–98 (9) TC</td>
<td></td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td>112–124 (17) C</td>
<td></td>
</tr>
<tr>
<td>Laboratory</td>
<td>190–210</td>
<td>67</td>
<td>87–97 (6)</td>
<td>3</td>
</tr>
<tr>
<td></td>
<td>220–235</td>
<td>67</td>
<td>102–110 (6)</td>
<td>3</td>
</tr>
<tr>
<td></td>
<td>220–235</td>
<td>240–270</td>
<td>95–103</td>
<td>4</td>
</tr>
</tbody>
</table>

Notes: Values in parentheses denote the average of the estimates’ standard errors. The expected \(\begin{array}{l}
\text{O}_3\text{ isotope composition} \\
on the VSMOW scale is calculated from the \(\text{O}_3\) enrichments reported relative to \(\text{O}_2\) using \(\delta^{18}\text{O}(\text{O}_3)_{\text{VSMOW}} = \\
\delta^{18}\text{O}(\text{O}_2)_{\text{VSMOW}} + 18\delta(\text{O}_3)_{\text{Air}-\text{O}_2} + [\delta^{18}\text{O}(\text{O}_2)_{\text{VSMOW}} \times 18\epsilon(\text{O}_3)_{\text{Air}-\text{O}_2}]\).
\end{array}\)

1 Observations (see Krankowsky et al. (2007) and refs. therein), lowermost values (19–25 km). Quoted temperature range is derived by matching measured \(\delta^{18}\text{O}(\text{O}_3)\) and laboratory data (see note 3).

2 This study, C1 observations (10–12 km). Letters denote the estimates derived using the data from Bhattacharya et al. (2008) and assuming only terminal (T), only central (C) and equiprobable terminal and central (TC) \(\text{O}_3\) atoms transfer to the artefact CO.

3 Calculated using the laboratory KIE temperature dependence data summarised by Janssen et al. (2003).

4 Calculated assuming a pressure dependence of the \(\text{O}_3\) formation KIE similar to that measured at 320 K (see Guenther et al. (1999) and refs. therein).
Figures
Fig. 1. (a) Distribution of CO mixing ratios as a function of concomitant O$_3$ mixing ratios measured by CARIBIC in the LMS ([O$_3$] > 300 nmol/mol). The shaded area is the two-dimensional histogram of the C2 measurements (all C2 data obtained until June 2013) counted in 5×1 nmol/mol size [O$_3$]×[CO] bins, thus darker areas emphasise greater numbers of particular CO–O$_3$ pairs observed. Small symbols denote the original C1 in situ measurements (black) and corrected for the artefacts (red); the C1 WAS analyses (11 of total 408) are shown with large symbols. Thin and thick step-lines demark the inner and outer statistical fences (ranges outside which the data points are considered mild or extreme outliers, see text) of the C2 data, respectively. The dashed curve exemplifies compositions expected from the linear mixing of very different (e.g., tropospheric and stratospheric) end-members. (b) Statistics on CO mixing ratios from C1 and C2 data shown in box-and-whisker diagrams for samples clustered in 20 nmol/mol O$_3$ bins (whiskers represent 9th/91st percentiles). (c) Sample statistic for each CARIBIC dataset (note the C2 figures scaled down by a factor of 1000). (d) Estimates of the C1 in situ CO contamination strength [CO$_c$] as a function of [O$_3$] (solid line) obtained by fitting the difference Δ[CO] between the C2 and C1 in situ [CO] (small symbols) as detailed in Appendix A (Eq. (A2)). Step line shows the Δ[CO] for the statistical averages (the shaded area equals the height of the inner statistical fences of the C2 data). Large symbols denote the estimates of [CO$_3$] in the C1 WAS data (slight variations vs. the in situ data are due to the sample mixing effects, see Sect. 3). Colour denotes the respective C1 WAS δ$^{18}$O(CO) (note that typically 6–7 in situ measurements correspond to one WAS sample).
Fig. 2. (accompanies Fig. 1) Carbon monoxide and ozone mixing ratios measured in C1. Small black symbols denote the C1 in situ measurements ($n = 12753$). The C1 WAS analyses ($n = 408$) are shown with large symbols; colour denotes the concomitant $\delta^{18}O$(CO) measurements. Thin and thick step-lines denote the inner and outer statistical fences of the C2 data, respectively. The dashed curve exemplifies compositions expected from the linear mixing of tropospheric and stratospheric end-members (see caption to Fig. 1 for details).
Fig. 3. $^{18}$O/$^{16}$O isotope composition of CO as a function of its reciprocal mixing ratio. Triangles present the data from the remote SH UT/LMS obtained by Brenninkmeijer et al. (1996) (B96). Colour refers to the concomitantly observed O$_3$ abundances; note the extremely low [O$_3$] encountered by B96 in the Antarctic "ozone hole" conditions. Filled and hollow circles denote the original and corrected (as exemplified by the dashed arrow) C1 WAS data, respectively, with the symbol size scaling proportional to the estimated contamination magnitude (see text).
Fig. 4. Measured C1 WAS $\delta^{18}$O(CO) (not corrected for artefacts) as a function of concomitant O$_3$ mixing ratio. Symbol colour denotes the artefact CO component (integral [CO$_c$] per each WAS); symbol size scales proportionally to the WAS CO mixing ratio corrected for artefacts (see Sect. 3 for details).
Fig. 5. Keeling plot of the data used in the calculations with the mixing model (MM). The C1 WAS isotope CO measurements are shown with symbols, solid lines denote the linear regressions through the various sets of samples selected by the MM ($n = 80$ sets are plotted). Colours refer to the $\delta^{13}$C (red) and $\delta^{18}$O (green) data, colour intensity indicates the coefficient of determination ($R^2$) of each regression, respectively. Darker colours denote higher $R^2$ values, with maxima of 0.92 for $\delta^{18}$O and 0.54 for $\delta^{13}$C data, respectively. The inferred contamination signatures $\delta$(CO$_c$) are found at ([CO]$^{-1}$ $\rightarrow$ 0. Regression uncertainties are shown in Fig. 6. Note that because different subsets of samples contain same data points, some of the symbols are plotted over (i.e., not all symbols contributing to a particular regression case may be seen).
Fig. 6. Results of the regression calculation with the MM. Shown with symbols are the contamination source isotope signatures $\delta$(CO$_c$) as a function of the respective coefficient of determination ($R^2$). Colour denotes the number of samples in each subset selected. Solid and dashed lines present the best guess ±1 standard deviation of the mean for the $\delta^{18}$O(CO$_c$) and $\delta^{13}$C(CO$_c$) estimates. Dashed circles mark the estimates obtained at highest $R^2$ for $\delta^{18}$O(CO$_c$) regression (above 0.9). See text for details.
Fig. 7. $^{18}$O/$^{16}$O and $^{13}$C/$^{12}$C isotope composition of CO measured in C1. Triangles present the data from the remote SH UT/LMS obtained by Brenninkmeijer et al. (1996) (B96). Colour refers to the concomitantly observed O$_3$ abundances; note the extremely low [O$_3$] encountered by B96 in the Antarctic ozone-hole conditions. Filled and hollow circles denote the original and corrected (as exemplified by the dashed arrow) C1 WAS data, respectively, with the symbol size scaling proportional to the estimated contamination magnitude (see text for details).