A theoretical study of abiotic methylation reactions of gaseous elemental mercury by halogen containing molecules

L. Castro¹, A. Dommergue², C. Larose³, C. Ferrari², and L. Maron¹

¹Université de Toulouse, INSA, UPS, LPCNO, 135 Avenue de Rangueil, 31077 Toulouse, France
²Laboratoire de Glaciologie et Géophysique de l’Environnement CNRS/Université Joseph Fourier Grenoble I, France
³Environmental Microbial Genomics Group, Laboratoire AMPEPER, UMR CNRS 5005, Ecole Centrale de Lyon, France

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Correspondence to: L. Maron (laurent.maron@irsamc.ups-tlse.fr)

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Abstract

Methylation reactions of gaseous elementary mercury by halogen containing molecules such as halogenomethane species CH₃X (with X=Cl, Br and I) and the dimethylchlorinium ion CH₃ClCH₃⁺ were investigated at the DFT level. With CH₃X, the reaction is predicted to be almost athermic and kinetically demanding for a thermal reaction. The reaction can proceed photochemically in the visible range; therefore sunlight may increase the reaction rate. These results compare well with the experimental data. Consecutive methylation of the CH₃HgX products (with X=Cl, Br and I) and subsequent formation of CH₃HgCH₃ were also studied. These reactions are predicted to be kinetically inaccessible and thermodynamically unfavorable. With CH₃ClCH₃⁺, the reaction is predicted to be athermic but kinetically easy. This is due to the suitability of the methyl transfer reagent. Geometrical and electronic data were systematically analyzed in order to rationalize the results.

1 Introduction

Methylmercury (MeHg) is a potent neurotoxin that provokes irreversible damage to the nervous system and can cause death. It is formed through the methylation of mercury (Hg), a toxic element found both naturally and as a human-introduced compound in the environment (Nriagu, 1988, 1989; Fitzgerald et al., 1998). Hg can be transported far from its emission sources, which include mainly fossil fuel combustion, industry and mining (AMAP/UNEP, 2008). It can be deposited after its oxidation to the more water soluble and reactive form divalent mercury (Hg(II)). Then, it has the potential to be converted into organo-metallic forms such as MeHg, which is likely the most ecologically relevant form. MeHg can bioaccumulate and biomagnify in the ecosystem and is found at high and increasing levels in marine species and freshwater systems all over the world including Arctic fauna (Wagemann et al., 1996). Of particular concern
are the high blood Hg concentrations observed in Northern people, whose traditional diet consists of marine organisms.

The mechanisms that produce MeHg are to date unknown, even though many pathways have been suggested (Barkay and Poulain, 2007). The biotic production of MeHg from inorganic mercury in aquatic sediments was first proposed by Jensen and Jernelov (1969). Biotic methylation was shown to be carried out by sulfate-reducing bacteria in anoxic sediments (Gilmour et al., 1992), and more recent evidence suggests that iron-reducing bacteria may methylate mercury as well (Fleming et al., 2006). Recent studies have shown that methylation likely occurs in oxic environments, although the processes are unknown (Cossa et al., 2009; Monperrus et al., 2007).

It has also been shown that the transfer of methyl groups can take place in abiotic conditions without the support of an enzymatic process. Chen et al. (2007) and Hammerschmidt et al. (2007) suggested that this process may be occurring in the aqueous phase in the atmosphere. Abiotic methylation of Hg(II) (Weber, 1993; Siciliano et al., 2005) requires suitable methyl donors and, although available methyl donors may be products of biological processes such as organic compounds (Kawamura et al., 1996), methylation is still considered to be abiotic (Craig, 1986; Falter and Wilken, 1998). Some potential pathways of chemical methylation have been studied, such as the reaction of Hg(II) with fulvic and humic acids (Nagase et al., 1982), acetic acid (Gardfeldt et al., 2003), methylmetals like methyltin or methyllead (Howell et al., 1986; Celo et al., 2006; Woggon et al., 1984; Cerrati et al., 1992), or small molecules such as methyl iodide (Celo et al., 2006; Ebinghaus and Wilken, 1993).

The contribution of abiotic methylation pathways may have been overlooked, in particular methylation involving Hg(0). Since Hg(0) is the dominant mercury species in the troposphere, abiotic Hg(0) methylation might be an important pathway by which MeHg species are produced and dispersed globally. To the best of our knowledge, only methyl iodide (CH₃I), a well known methylating agent in organometallic chemical synthesis, has been proposed as a possible candidate for Hg(0) methylation (Maynard, 1932; Weber, 1993; Hall et al., 1995; Tokos et al., 1998; Celo et al., 2006).

However, experimental studies have shown important disagreement. Maynard was the first to observe Hg(0) methylation by CH₃I at room temperature in the sunlight (Maynard, 1932), but the yield was found to be low and the reaction rate slow. Weber (1993) also proposed the possible formation of CH₃HgI by the same reaction. In a series of experiments, Hall et al. (1995) reported the absence of Hg(0) methylation in the aqueous phase in the dark, and observed an oxidation of Hg(0) in full sunlight. Direct methylation of Hg(0) by dimethyl sulfide (DMS) was not observed, but slow formation of methylmercury close to the detection limit of the apparatus was observed using Hg(0) and CH₃I (Hall et al., 1995). The authors concluded that Hg(0) methylation by CH₃I was not significant for the cycle of mercury. This was then confirmed by experiments conducted by Tokos et al. (1998). Finally, Celo et al. (2006) found that methylation occurs in the dark at a low rate, but with a yield equivalent to that reported for methylation by sulfate-reducing bacteria.

Theoretical studies, notably the DFT (Density Functional Theory) method, can be very useful in reconciling experimental inconsistencies since they are able to reliably reproduce the reactivity of organometallic systems (Chemical Review, 2000). These kinds of treatments have already been used for two precedent works involving mercury: its reaction with halogen radicals and anions (Maron et al., 2008) and its reactivity with ozone (Castro et al., 2009). Other theoretical works are present in the literature (e.g., Tossell, 2003; Strömberg et al., 1991; Shepler and Peterson, 2003; Hu et al., 1999; Kaupp et al., 1994), but there is a lack of research in the domain. Indeed, field observations and laboratory experiments can be supported by theoretical studies and insights brought by these methods should not be neglected. In this paper, several reactions that could potentially lead to the methylation of Hg(0) were investigated in the gas phase. Halogenomethane species (CH₃X with X=Cl, Br, I) were selected as potential candidates for mercury methylation not only because they are widely distributed in the marine boundary layer and in the free troposphere (Cox et al., 2003, 2005), but also to enable comparison with the experimental studies involving CH₃I. The different formation reactions of mono-methyl and di-methyl mercury from
Hg(0) and halogenomethane species were studied. In addition, the reaction between Hg(0) and the dimethylchlorinium ion CH₃Cl⁺, which may potentially lead directly to HgCH₃⁺, was also examined.

2 Computational details

Mercury (Kuechle et al., 1991) and all the halogen atoms (Bergner et al., 1993) were treated with Stuttgart-Dresden pseudopotentials in combination with their adapted basis set, augmented by a set of polarization functions (f for Hg and d for halogens) (Ehlers et al., 1993). Carbon and hydrogen atoms were described with a 6-31G(d,p) double-ζ basis set (Hehre et al., 1972). Calculations were carried out at the DFT level of theory using the hybrid functional B3PW91 (Perdew and Wang, 1992; Becke, 1993). Geometry optimizations were carried out without any symmetry restrictions, and the nature of the extrema (minima and transition states) was verified with analytical frequency calculations. Gibbs free energies were obtained at T = 298.15 K within the harmonic approximation. IRC calculations were performed to check the connections of the optimized transition states. DFT calculations were carried out with the Gaussian03 suite of program (Gaussian 03 (Revision D.02) et al., 2004). Finally, the electronic density (at the DFT level) was analyzed using the Natural Bond Orbital (NBO) technique (Reed et al., 1988).

3 Results and discussion

3.1 Reaction of Hg(0) with CH₃X

Reactions between Hg(0) and CH₃X (with X=Cl, Br, I) were computed (Fig. 1). The reaction is an oxidative addition of CH₃X to Hg(0) in order to form the divalent complex CH₃Hg(II)X. The geometries are analogous for the three halogens, but bond distances vary because of the increase of the halogen atomic radius with the atomic number. Thus, the C–X bond in CH₃X is longer for X=I (2.17 Å) than for X=Br (1.95 Å) and X=Cl (1.80 Å). The same behaviour is found for the Hg–X bond length in the reaction product (2.66 Å for X=I, 2.47 Å for X=Br and 2.34 Å for X=Cl). The reaction is predicted to be almost athermic for the three halogens. Indeed, even if it is slightly endergonic for X=Cl and X=Br, and slightly exergonic for X=I, it is not possible to conclude since the values are within the systematic error of the DFT method (estimated at 5 kcal mol⁻¹) (Schultz et al., 2008; Zhao and Truhlar, 2008). These thermodynamics are due to the fact that a strong C–X bond is broken during the reaction while two weaker bonds (Hg–C and Hg–X) are formed. The loss and gain of energies reach a balance and induce a globally athermic reaction. Nevertheless, CH₃HgX formation is slightly more favorable when the atomic number of the halogen increases. This is due to the lower C–X bond strength in the CH₃X reactant. Indeed, the electronegativity of X decreases when the atomic number increases, and therefore the C–X bond polarity and strength decrease as well. In any event, these energetic values are characteristic of feasible reactions for the three halogens from a thermodynamic point of view, and even slightly favorable for X=I.

The calculated activation barriers are found to be high (+67.6 kcal mol⁻¹ for X=Cl, +64.8 kcal mol⁻¹ for X=Br and +60.4 kcal mol⁻¹ for X=I), indicating kinetically demanding reactions. The height of the barriers can be explained by geometrical and charge analyses at the transition state level (Fig. 2). The structure is a 3-centered triangle pattern characteristic of oxidative additions, where the methyl group is transferred from X to Hg. Geometrically, it is noteworthy that all the interatomic distances are longer than the corresponding ones in both reactants and products. Indeed, the Hg–C distance is equal to 2.51 Å whereas it is almost equal to 2.09 Å in CH₃HgX. The Hg–X ones are longer too, since the Hg–X bond lengths in CH₃HgX are 2.34 Å, 2.47 Å and 2.66 Å for X=Cl, Br and I, respectively. Finally, the C–X distances in CH₃X are 1.80 Å, 1.95 Å and 2.17 Å for X=Cl, Br and I, respectively, which is shorter than in the transition state structures. Thus, all bonds are not yet fully broken or formed at the transition state.
NBO analysis was performed at the transition state. The charges of the halogen atoms are equivalent to the ones found in the CH$_3$HgX product. The halogens attract the electronic density from the mercury and from the methyl group in order to be sufficiently charged to form the Hg–X bond. The methyl group is positively charged, which induces an electrostatic repulsion between Hg and CH$_3$. However, the methyl group is negatively charged in the CH$_3$HgX product (–0.30 for the three halogens), so an important electronic reorganization is needed. From the second-order donor-acceptor analysis, it is noteworthy that the 3p (4p) orbitals of Cl (Br) give some electronic density not only to a 6p orbital of Hg, which forms the Hg–Cl (Hg–Br) bond, but also to the antibonding molecular orbital of the Hg–C bond, in order to break it. This is consistent with a slightly late transition state. For X=I, at the second order, a 5p orbital of I gives electronic density to a 6p orbital of Hg, and the 6s orbital of Hg donates to the C–I antibonding molecular orbital. The transition state is no longer late for X=I, thus in total agreement with the Hammond postulate, stating that transition states are late for endergonic reactions (Hammond, 1955). The activation barrier is lower when the atomic number increases. This is also related to the C–X bond strength in the CH$_3$X reactant. This reaction is difficult from a kinetic point of view, but sunlight irradiation can help the reaction to occur photochemically. Indeed, the absorption radiations needed in order to cross the barrier are equal to 457 nm, 479 nm and 518 nm for X=Cl, Br and I, respectively. These are radiations in the visible range; therefore these reactions may be slightly faster and observable in the presence of sunlight.

Some interesting links appear between this theoretical study and the experimental ones present in the literature. Indeed, calculations predict a thermodynamically favorable reaction between Hg(0) and CH$_3$I. This is in agreement with the experimental results by Celo et al. (2006) where MeHg formation was shown to occur in the dark. However, the activation barrier is predicted to be high, which is characteristic of a kinetically difficult reaction. It is for this reason that the formation of methylmercury observed by Celo et al. (2006) is slow. Moreover, Hall et al. (1995) were likely unable to identify any methylation in the dark or in the aqueous phase due to the slowness of the reaction.

On the other hand, the theoretical study shows that the reaction should be faster with the help of solar irradiation. This is in agreement with Hall et al. (1995) who reported the formation of a small amount of methylmercury from gas phase experiments in full sunlight. The reaction was very slow in these conditions, which is consistent with the high activation barrier, but may also be due to other competitive reactions such as the photodissociation of CH$_3$I.

The products of the reactions are CH$_3$HgX. We then determined whether it is possible to photodissociate these complexes to form either HgCH$_3$ or HgCH$_3^+$ (Reactions 1 and 2).

\[ \text{CH}_3\text{HgX} \xrightarrow{\text{hv}} \text{CH}_3\text{Hg}^+ + \text{X}^- \] (R1)

\[ \text{CH}_3\text{HgX} \xrightarrow{\text{hv}} \text{CH}_3\text{Hg} + \text{X} \] (R2)

\[ \text{CH}_3\text{HgX} \xrightarrow{\text{hv}} \text{CH}_3 + \text{HgX} \] (R3)

For dissociation (1), wavelengths of 167 nm, 170 nm and 176 nm are needed for X=Cl, Br and I, respectively. These wavelengths are filtered in the upper atmosphere in the troposphere; therefore the tropospheric formation of HgCH$_3^+$ is not possible. For dissociation (2), radiations of 329 nm, 367 nm and 413 nm are needed for X=Cl, Br and I, respectively. In theory, it is thus possible to form HgCH$_3$ and X from Hg(0) and CH$_3$I in the presence of sunlight in the troposphere. This photodissociation is however in competition with the photoformation of HgX (3) for which wavelengths of 437 nm, 457 nm and 485 nm are needed for X=Cl, Br and I, respectively. Dissociations (2) and (3) are in competition, thus significantly reducing the formation of HgCH$_3$ from CH$_3$I.

### 3.2 Reactions of dimethylmercury formation from CH$_3$HgX

In this section, two possible pathways of dimethylmercury formation from CH$_3$HgX were studied. First, reactions between CH$_3$HgX and another molecule of CH$_3$X to form...
CH₂HgCH₃ and X₂ were computed (Fig. 3). The geometries are analogous for the three halogens.

The reactions are found to be very endergonic (+58.4 kcal mol⁻¹ for X=Cl, +45.8 kcal mol⁻¹ for X=Br and +37.3 kcal mol⁻¹ for X=I). This is due to an important bond energy difference between C–X and Hg–X bonds and the formed X–X and Hg–C bonds. CH₃HgCH₃ is not a particularly stable product, in agreement with experimental Hg–C bond dissociation energy (57 kcal mol⁻¹) (Handbook of Chemistry and Physics (89th edition), 2009). The free energy of reaction is higher when the atomic number decreases for the same reason as in the previous reaction. From a kinetic point of view, activation barriers are very high for the three halogens (+82.4 kcal mol⁻¹ for X=Cl, +70.0 kcal mol⁻¹ for X=Br and +61.2 kcal mol⁻¹ for X=I). The structure of the transition state is a metathesis one where a X atom transfers from a methyl group to the X ligand of CH₃HgX. The geometrical configuration is very distorted since the methyl group and both X atoms are almost aligned, which is classical of d² metathesis reactions (Maron and Eisenstein, 2001). The electronic pattern of metathesis transition state is introduced (Fig. 4). The B center has to be positively charged in order to stabilize the transition state thanks to the charge alternation (Maron et al., 2003). In the present case, one of the halogen atoms has to be positively charged. Thus, it is natural that the transition states lie at high energies since halogen atoms are rather electronegative.

Geometrically, only the inactive Hg–CH₃ bond is still formed since it remains constant throughout the reaction. The remaining structure does not exhibit any fully formed chemical bond, like in the case of the transition state of the reaction between Hg(0) and CH₃X (Fig. 5). Indeed, X–X bonds are longer than in free molecules (2.04 Å for Cl₂, 2.32 Å for Br₂ and 2.80 Å for I₂), as well as the Hg–C bond (2.11 Å in CH₃HgCH₃) and the C–X bonds (introduced earlier).

From NBO analysis, it is noteworthy that the most negatively charged halogen is the one which is away from the metal. Thus, no electrostatic stabilization is found between this halogen and Hg. For X=Cl and X=Br, electrostatic repulsions are observable between the two X atoms and between both CH₃X moieties. For X=I, the central halogen is positively charged, which brings slightly more stabilization, as can be seen from the energetic data (the activation barrier is lower for X=I). The C–X strength in CH₃X (as seen previously) is also important for the stabilization of the transition state when the atomic number of X increases.

Therefore, based on our calculations, this reaction is definitively not possible thermally and even photochemically, because thermodynamics push the reaction towards the formation of the reactants instead of the formation of the products.

Another reaction is conceivable in order to form dimethylmercury from CH₃HgX:

2CH₃HgX → CH₃HgCH₃ + HgX₂

This ligand exchange has already been observed between two CH₃Zn complexes to form CH₃ZnCH₃ and ZnI₂, so it is interesting to know if this reaction is possible with CH₃HgI, since Hg and Zn are both group 12 transition metals. The reaction was investigated with the three halogens (Fig. 6) and their geometries are analogous.

Thermodynamically, the reaction is endergonic for the three halogens. This is due to the low stability of CH₃HgCH₃. Indeed, the variation of entropy is insignificant in this case and HgX₂ is a very stable molecule. Kinetically, barriers are rather high. Geometrically, the structure is a typical ligand exchange square pattern. The two inactive Hg–X and Hg–C bond lengths remain the same as in CH₃HgCH₃ and HgX₂. The distance between the two Hg atoms is quite small (2.92 Å for X=Cl, 2.90 Å for X=Br and 2.85 Å for X=I) and induces the destabilization of the transition state. Indeed, a strong electrostatic repulsion was found between the two positively charged metals (Table 1). This repulsion is lower when the atomic number of X increases because the electronegativity of X decreases, therefore the Hg atoms are less positively charged (Table 1). The barriers vary in the same way.

The reaction is found to be kinetically difficult and thermodynamically unfavorable. It is thus unexpected that any formation of CH₃HgCH₃ occurs from CH₃HgX.
3.3 Reactions between Hg(0) and the dimethylchlorinium ion

The reaction between Hg(0) and the dimethylchlorinium ion CH₃ClCH₃⁺ was investigated because it is an interesting way to form HgCH₃⁺ directly (Fig. 7). The reaction consists of a CH₃⁺ transfer from the ion to the mercury, in order to form CH₂Cl and HgCH₃⁺. Thermodynamically, the reaction is found to be almost athermic since the systematic error of DFT is 5 kcal mol⁻¹. This is characteristic of a feasible reaction. The transition state structure is typical of a Sₘ2 nucleophilic substitution where the transferred methyl group is planar (Fig. 8). This is in agreement with the expected transfer of a CH₃⁺ group.

The activation barrier is very low, because CH₃ClCH₃⁺ is a very reactive species, which is known as a good CH₃⁺ donor. The difficult oxidation of Hg(0) observed in the previous reactions with CH₃X is counterbalanced by the breaking of the weak Cl–C bond of the dimethylchlorinium ion. From a geometrical point of view, the transition state is intermediate between reactants and products. Indeed, the inactive Cl–C bond length is 1.82 Å, intermediate between the Cl–C ones in CH₂ClCH₃⁺ (1.85 Å) and in CH₂Cl (1.80 Å). The other Cl–C bond length is elongated up to 2.34 Å. The Hg–C bond length is 2.54 Å whereas in HgCH₃⁺ it is 2.16 Å. From NBO analysis, the Cl atom is less positively charged in the transition state (+0.19) than in the reactant (+0.50), which is very favorable since Cl is an electronegative element. Moreover, both the CH₃–Cl and Hg–C bonds are already formed. From the second-order donor-acceptor analysis, the 3p orbital of the Cl atom gives electronic density to the antibonding orbital of the Hg–C bond. This reaction is thus possible, even if it is not pushed towards the formation of the products. This is a demonstration of the importance of the nature of the methyl transfer agent. It has to be a reactive molecule that can transfer its methyl group to become more stable, like the dimethylchlorinium ion, in order to compensate for the difficulty of mercury oxidation. This seems easier when the transferred methyl group is planar and positively charged.

4 Conclusions

In this paper, theoretical studies were used for two different purposes. First, experimental results concerning the reaction of Hg(0) with CH₃I were rationalized. This reaction is predicted to be feasible, slightly exergonic, but with a high activation barrier. It is thus kinetically difficult, which is in agreement with its poor observation (Hall et al., 1995) and with the low rate observed experimentally (Celo et al., 2006; Maynard, 1932). Solar irradiation may accelerate the rate of the reaction, as observed by Hall et al. The second approach aims at gaining some insights into the chemical parameters which control the Hg(0) methylation. Some possible reactions of Hg(0) with different substrates in order to form HgCH₃ or HgCH₃⁺ were investigated. The oxidation and methylation of Hg(0) are possible only if the methyl transfer agent is well suited. Indeed, reactions with CH₃X are globally athermic and kinetically difficult because of the high stability of these molecules. Moreover, the transfer of a neutral methyl group seems to be quite difficult. On the other hand, the reaction with the dimethylchlorinium ion CH₂ClCH₃⁺ is thermodynamically feasible and kinetically easy. This is associated with the reactive nature of this molecule due to the weak Cl–C bond and with the transfer of a positively charged methyl group. Finally, the formation of dimethylmercury is not possible with CH₃X or another molecule of CH₂HgX. This is mainly due to electrostatic repulsions at the transition state level.

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Table 1. Charges of mercury atoms in the transition state structures of the reaction between two CH₃HgX complexes, from a NBO analysis. Hg₁ is the Hg–X inactive bond, Hg₂ is the Hg–C inactive bond.

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<th>Cl</th>
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<tr>
<td>Hg₁</td>
<td>+0.79</td>
<td>+0.64</td>
<td>+0.42</td>
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<tr>
<td>Hg₂</td>
<td>+0.82</td>
<td>+0.79</td>
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longer for X=I (2.17 Å) than for X=Br (1.95 Å) and X=Cl (1.80 Å). The same behaviour is found for the Hg-X bond length in the reaction product (2.66 Å for X=I, 2.47 Å for X=Br and 2.34 Å for X=Cl). The reaction is predicted to be almost at hermic for the three halogens. Indeed, even if it is slightly endergonic for X=Cl and X=Br, and slightly exergonic for X=I, it is not possible to conclude since the values are within the systematic error of the DFT method (estimated at 5 kcal.mol$^{-1}$) (Schultz et al., 2008; Zhao and Truhlar, 2008).

These thermodynamics are due to the fact that a strong C-X bond is broken during the reaction while two weaker bonds (Hg-C and Hg-X) are formed. The loss and gain of energy reach a balance and induce a globally athermic reaction. Nevertheless, CH$_3$HgX formation is slightly more favourable when the atomic number of the halogen increases. This is due to the lower C-X bond strength in the CH$_3$X reactant. Indeed, the electronegativity of X decreases when the atomic number increases, and therefore the C-X bond polarity and strength decrease as well.

In any event, these energetic values are characteristic of feasible reactions for the three halogens from a thermodynamic point of view, and even slightly favorable for X=I.

Fig. 1. Calculated Gibbs free energy profiles for the reaction of Hg(0) with CH$_3$X. The 3-D structures are the optimized structures for X=Cl.

Fig. 2. Bond distances (black values) and charges (red values) at the transition state level of the reaction of Hg(0) with CH$_3$X.
Fig. 3. Calculated Gibbs free energy profiles of the reaction of \( \text{CH}_3\text{HgX} \) with \( \text{CH}_3\text{X} \). The 3-D structures are the optimized structures for \( X=\text{Cl} \).

\[ \Delta G^\circ \text{(neutral)} \]

\[
\begin{array}{c}
\text{A} & \text{M} & \text{R} & \text{B} \\
\delta^- & & & \delta^+ \\
\text{B} & \text{M} & \text{R} & \text{A} \\
\delta^- & & & \delta^+ \\
\end{array}
\]

Fig. 4. Metathesis transition state pattern.
2.80 Å for I$_2$, as well as the Hg-C bond (2.11 Å in CH$_3$HgCH$_3$) and the C-X bonds (introduced earlier).

**Figure 5.** Bond distances (right) and charges (left) at the transition state level of the reaction of CH$_3$HgX with CH$_3$X.

From NBO analysis, it is noteworthy that the most negatively charged halogen is the one away from the metal. Thus, no electrostatic stabilization is found between this halogen and Hg. For X=Cl and X=Br, electrostatic repulsions are observable between the two X atoms and between both CH$_3$X moieties. For X=I, the central halogen is positively charged, which brings slightly more stabilization, as can be seen from the energetic data (the activation barrier is lower for X=I). The C-X strength in CH$_3$X (as seen previously) is also important for the stabilization of the transition state when the atomic number of X increases.

Therefore, based on our calculations, this reaction is definitively not possible thermically and even photochemically, because thermodynamics push the reaction towards the formation of the reactants instead of the formation of the products.

Another reaction is conceivable in order to form dimethylmercury from CH$_3$HgX:

This ligand exchange has already been observed between two CH$_3$ZnI complexes to form CH$_3$ZnCH$_3$ and ZnI$_2$, so it is interesting to know if this reaction is possible with CH$_3$HgI, since Hg and Zn are both group 12 transition metals.

The reaction was investigated with the three halogens (figure 6) and the geometries are analogous.

**Figure 6.** Calculated Gibbs free energy profiles of the reaction between two CH$_3$HgX complexes. The 3-D structures are the optimized structures for X=Cl.

**Fig. 5.** Bond distances (right) and charges (left) at the transition state level of the reaction of CH$_3$HgX with CH$_3$X.

**Fig. 6.** Calculated Gibbs free energy profiles of the reaction between two CH$_3$HgX complexes. The 3-D structures are the optimized structures for X=Cl.
The reaction consists of a CH$_3^+$ transfer from the ion to the mercury, in order to form CH$_3$Cl and HgCH$_3^+$. Thermodynamically, the reaction is found to be almost athermic since the systematic error of DFT is 5 kcal.mol$^{-1}$. This is characteristic of a feasible reaction. The transition state structure is typical of a S$_N$2 nucleophilic substitution where the transferred methyl group is planar (figure 8). This is in agreement with the expected transfer of a CH$_3^+$ group.

**Fig. 7.** Calculated Gibbs free energy profile of the reaction of Hg(0) with CH$_3$ClCH$_3^+$.

**Fig. 8.** Bond distances (black values) and charges (red values) at the transition state level of the reaction between Hg(0) and CH$_3$ClCH$_3^+$. 